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Question 2 $2 \text{ Fe(s)} + 3 \text{ O}_2\text{(g)} \rightarrow \text{Fe}_2\text{O}_3\text{(s)}$ $\Delta H_f^\circ = -824 \text{ kJ mol}^{-1}$

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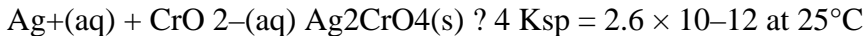
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Answer the following questions relating to the solubilities of two silver compounds, Ag_2CrO_4 and Ag_3PO_4 . Silver chromate dissociates in water according to the equation shown below. ? 2

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$\text{CH}_3\text{CH}_2\text{COOH}(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{CH}_3\text{CH}_2\text{COO}^-(\text{aq}) + \text{H}^+$

$\text{O}^+(\text{aq})$ Propanoic acid, $\text{CH}_3\text{CH}_2\text{COOH}$, is a carboxylic acid that reacts with water according to the equation above. At 25 C the pH of a 50.0 mL sample of 0.20 M $\text{CH}_3\text{CH}_2\text{COOH}$

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$\text{Ag}^+(\text{aq}) + \text{CrO}_4^{2-}(\text{aq}) \rightleftharpoons \text{Ag}_2\text{CrO}_4(\text{s})$? 4 $K_{sp} = 2.6 \times 10^{-12}$ at 25°C

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Answer the following questions about carbon monoxide, CO(g), and carbon dioxide, CO₂(g). Assume that both gases exhibit ideal behavior. (a) Draw the complete Lewis structure (electron-dot

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diagram) for the CO molecule and for the CO₂ molecule. 1 point for each correct, complete Lewis structure

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